

## Redox behaviour, electrochromic properties and photoluminescence of potassium lanthano phosphomolybdate sandwich-type compounds

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### Supporting Information

#### Captions to Figures

**Figure S1.** Infrared spectra in the range: A) 4000-500  $\text{cm}^{-1}$  and B) 1300-600  $\text{cm}^{-1}$  for potassium salts: (a)  $\text{Sm}(\text{PMo}_{11})_2$ , (b)  $\text{Eu}(\text{PMo}_{11})_2$ , (c)  $\text{Gd}(\text{PMo}_{11})_2$ , (d)  $\text{Tb}(\text{PMo}_{11})_2$  and (e)  $\text{Dy}(\text{PMo}_{11})_2$ .

**Figure S2.**  $^{31}\text{P}$  NMR spectra in  $\text{D}_2\text{O}$  of: (A)  $\text{Sm}(\text{PMo}_{11})_2$ , (B)  $\text{Eu}(\text{PMo}_{11})_2$ , (C)  $\text{Gd}(\text{PMo}_{11})_2$ , (D)  $\text{Tb}(\text{PMo}_{11})_2$  and (E)  $\text{Dy}(\text{PMo}_{11})_2$ .

**Figure S3. A)** Cyclic voltammograms of  $\text{K}^+$  salt of  $\text{Eu}(\text{PMo}_{11})_2$  ( $5 \times 10^{-4} \text{ mol dm}^{-3}$ ) in pH 3.0  $\text{H}_2\text{SO}_4/\text{Na}_2\text{SO}_4$  buffer solution at scan rates of 0.02, 0.04, 0.06, 0.08, 0.1, 0.15, 0.2, 0.25, 0.3, 0.35, 0.4, 0.45 and  $0.5 \text{ V s}^{-1}$ .

**B)** Cyclic voltammograms of  $\text{K}^+$  salt of  $\text{Gd}(\text{PMo}_{11})_2$  ( $5 \times 10^{-4} \text{ mol dm}^{-3}$ ) in pH 3.0  $\text{H}_2\text{SO}_4/\text{Na}_2\text{SO}_4$  buffer solution at scan rates of 0.02, 0.04, 0.06, 0.08, 0.1, 0.15, 0.2, 0.25, 0.3, 0.35, 0.4, 0.45 and  $0.5 \text{ V s}^{-1}$ .

**C)** Cyclic voltammograms of  $\text{K}^+$  salt of  $\text{Tb}(\text{PMo}_{11})_2$  ( $5 \times 10^{-4} \text{ mol dm}^{-3}$ ) in pH 3.0  $\text{H}_2\text{SO}_4/\text{Na}_2\text{SO}_4$  buffer solution at scan rates of 0.02, 0.04, 0.06, 0.08, 0.1, 0.15, 0.2, 0.25, 0.3, 0.35, 0.4, 0.45 and  $0.5 \text{ V s}^{-1}$ .

**D)** Cyclic voltammograms of  $\text{K}^+$  salt of  $\text{Dy}(\text{PMo}_{11})_2$  ( $5 \times 10^{-4} \text{ mol dm}^{-3}$ ) in pH 3.0  $\text{H}_2\text{SO}_4/\text{Na}_2\text{SO}_4$  buffer solution at scan rates of 0.02, 0.04, 0.06, 0.08, 0.1, 0.15, 0.2, 0.25, 0.3, 0.35, 0.4, 0.45 and  $0.5 \text{ V s}^{-1}$ .

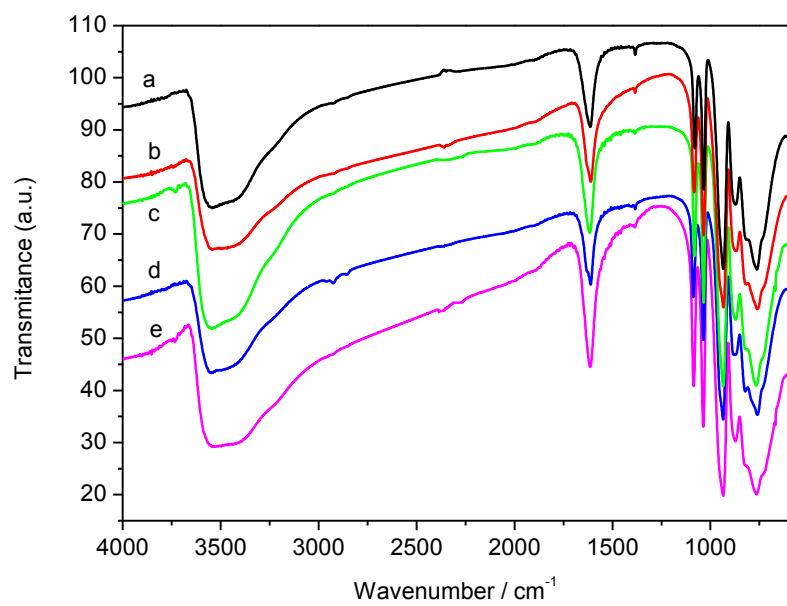
**Figure S4.** UV-visible spectra of  $\text{Sm}(\text{PMo}_{11})_2$  (A),  $\text{Eu}(\text{PMo}_{11})_2$  (C),  $\text{Gd}(\text{PMo}_{11})_2$  (E) and  $\text{Dy}(\text{PMo}_{11})_2$  salts (G) in pH 3.0  $\text{H}_2\text{SO}_4/\text{Na}_2\text{SO}_4$  buffer solution before

(a) and after reduction at 0.1 V for 30 min to 6.0 h (b to i); Absorbance evolution versus reduction time (B, D, F and H).

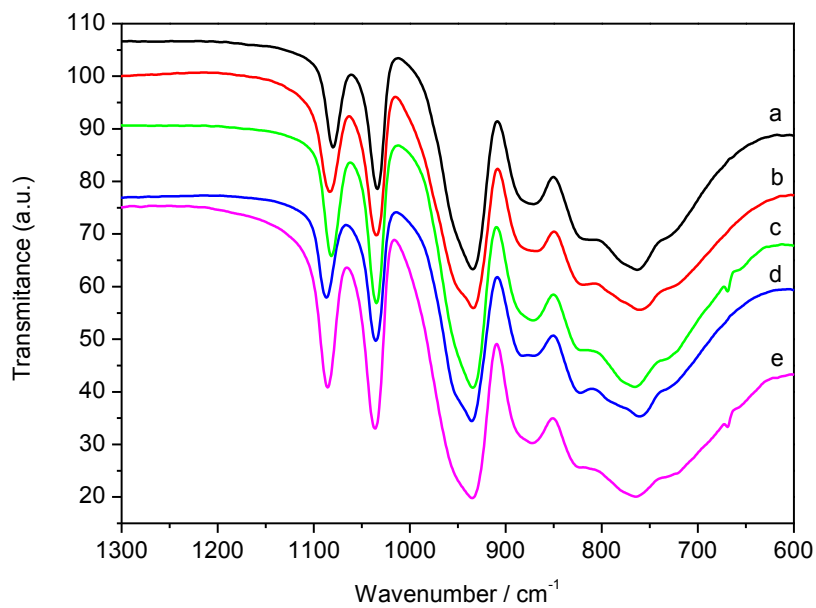
**Figure S5.** Emission decay curves acquired at (A) 10 K and (B) 300 K of EuPOM monitored at 614 nm and excited at 465 nm. The solid lines correspond to the data best fit using a single exponential function. The insets show the respective regular residual plots and the  $\chi^2_{\text{red}}$  values for a better judgment of the fit quality.

**Figure S1.**

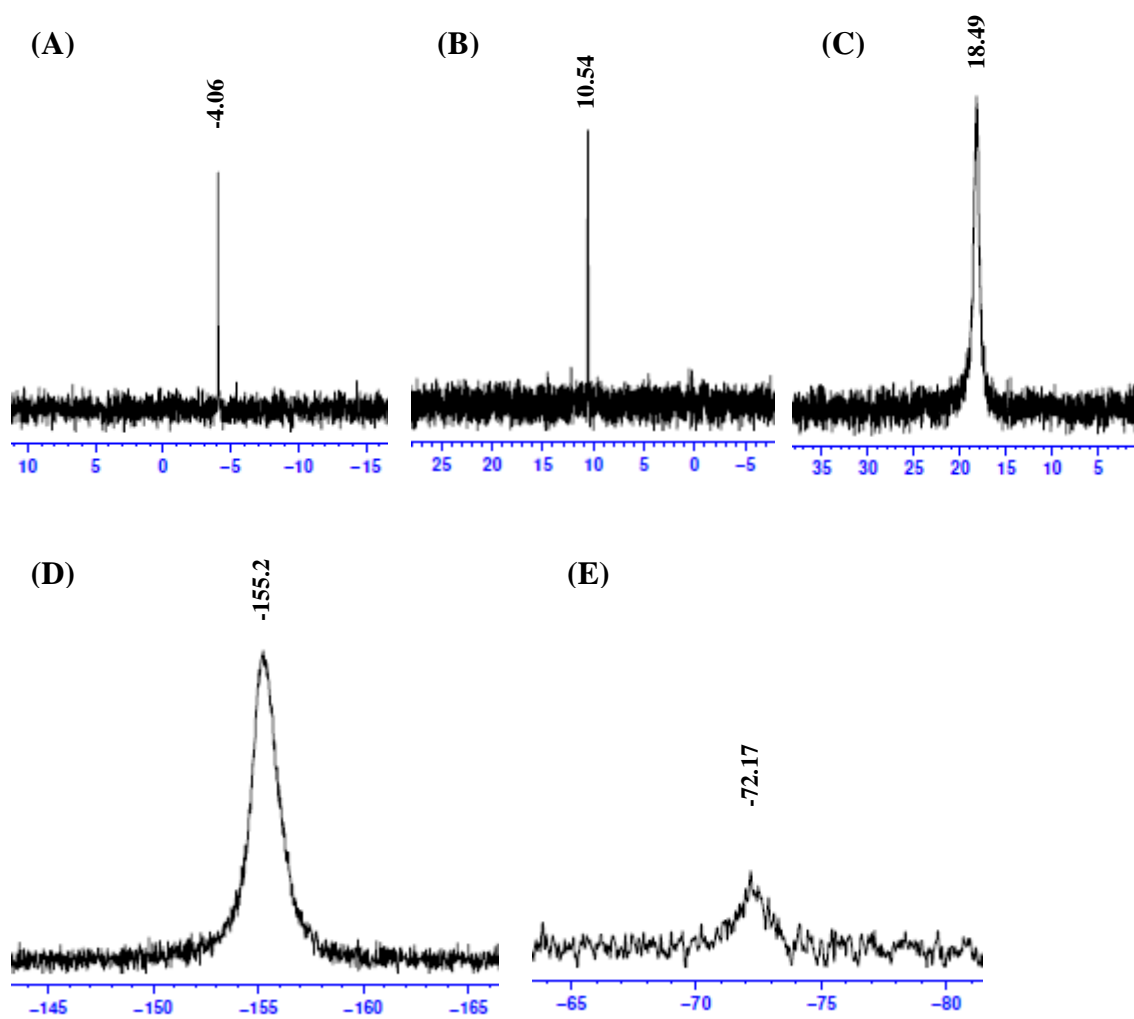
**(A)**



**(B)**

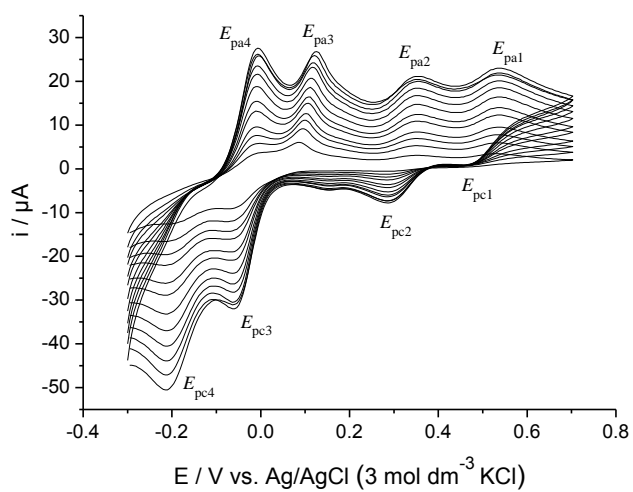


**Figure S2.**

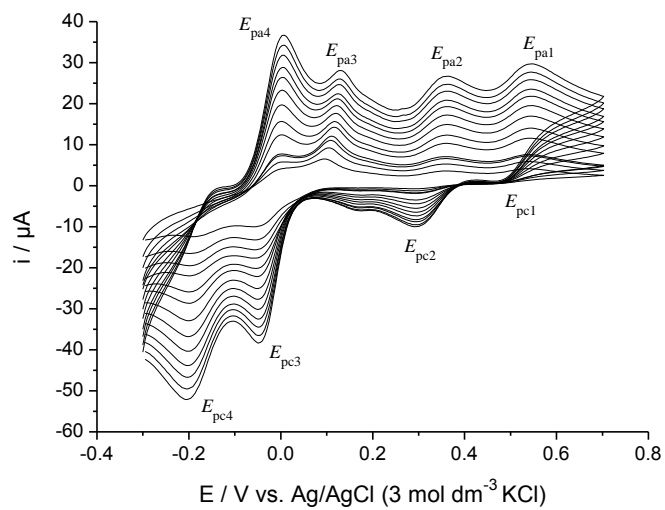


**Figure S3.**

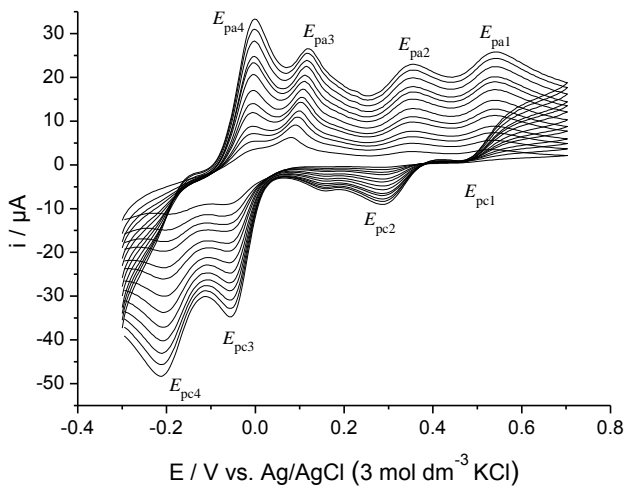
**(A)**

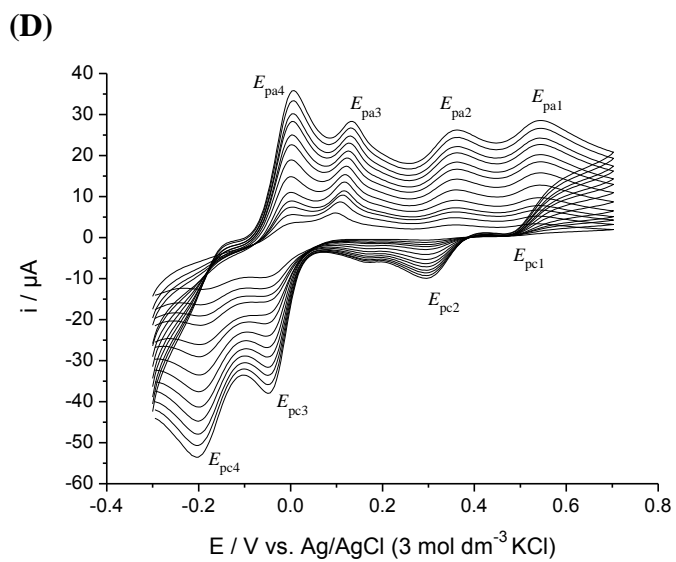


**(B)**

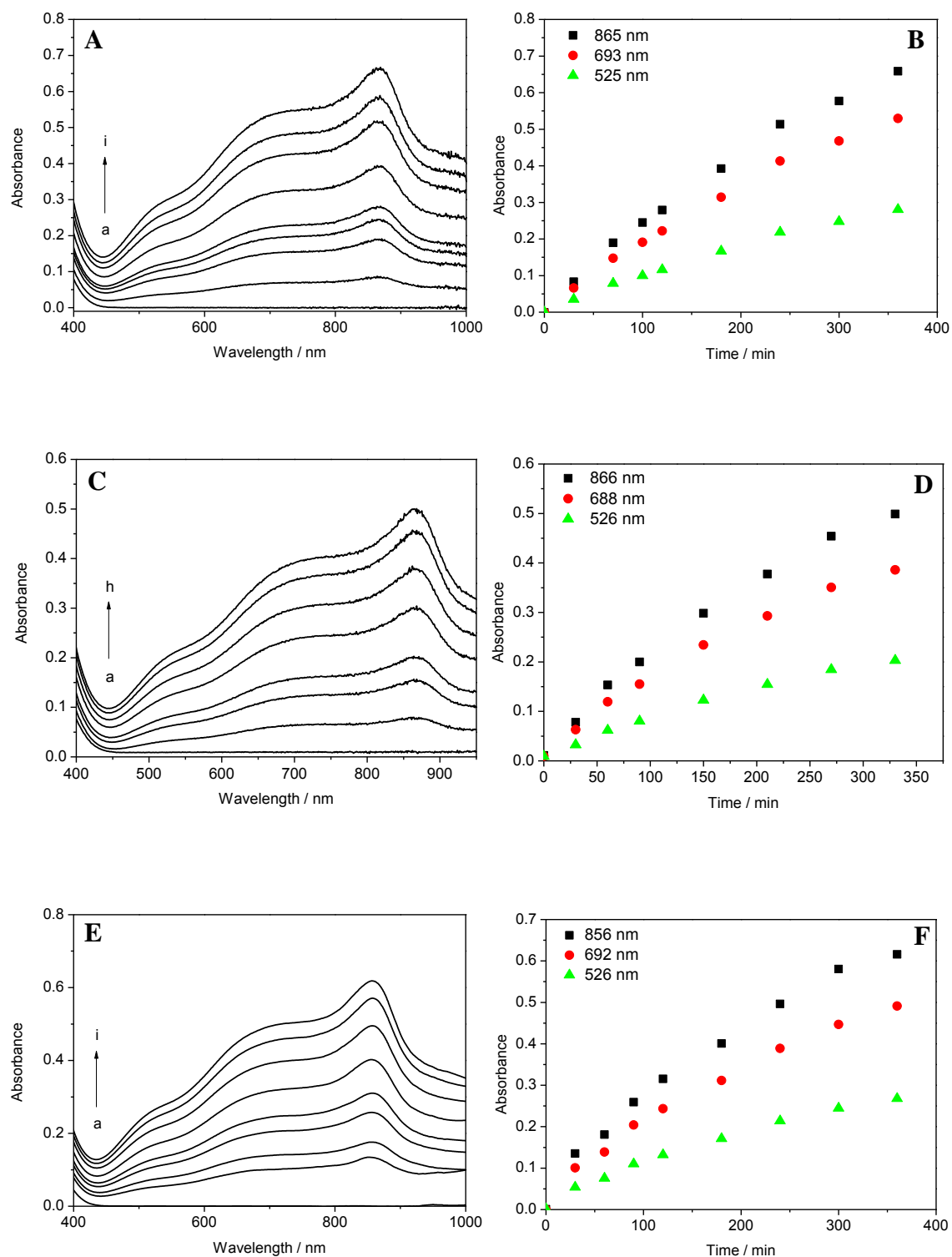


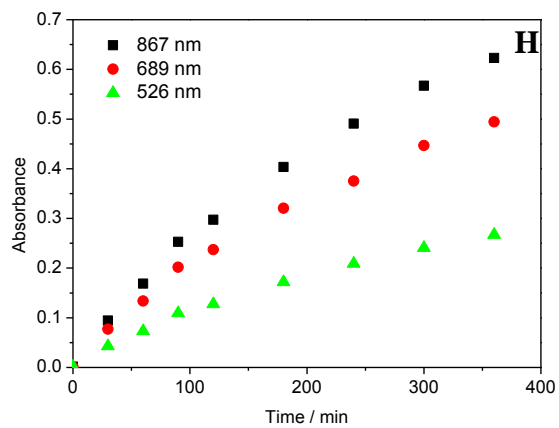
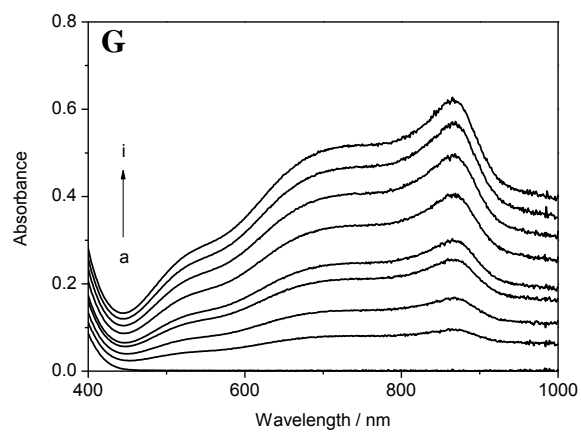
**(C)**





**Figure S4.**







**Figure S5.**

