

Electrochemically activated graphite electrode with excellent kinetics of electrode processes of V(II)/V(III) and V(IV)/V(V) couples in a vanadium redox flow battery

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Supporting Materials

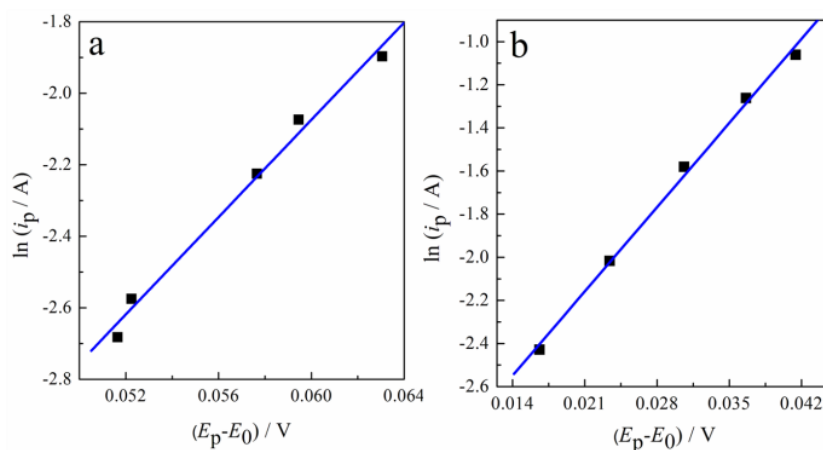


Fig. S1. The plots of $\ln(i_p)$ versus $(E_p - E_0)$ for the oxidation of V(IV) on the pristine graphite electrode (a) and EAGE (b) in $2 \text{ mol} \cdot \text{dm}^{-3} \text{ H}_2\text{SO}_4 + 2 \text{ mol} \cdot \text{dm}^{-3} \text{ VOSO}_4$ solution

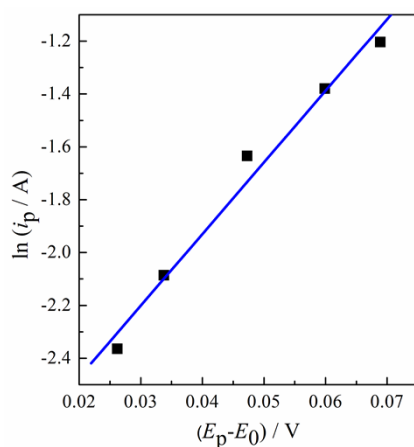


Fig. S2. The plot of $\ln(i_p)$ versus $(E_p - E_0)$ for the reduction of V(III) on EAGE in $2 \text{ mol} \cdot \text{dm}^{-3} \text{ H}_2\text{SO}_4 + 2 \text{ mol} \cdot \text{dm}^{-3} \text{ V(III)}$ solution