Supporting Information

Enhancement of Mineralization Ability of C₃N₄ via lower Valence Position by Tetracyanoquinodimethane Organic Semiconductor

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Figure S1 TEM images of pure $g-C_3N_4$ (a) TCNQ- C_3N_4 with TCNQ mass fraction as 1 %, 5 %, 10 %, 20% (b-e) and pure TCNQ (f)



Figure S2 (a) Photocatalytic degradation of phenol and (b) the apparent rate constants over pure g-C₃N₄ (0%), pure TCNQ (100%) and TCNQ-C₃N₄ with different TCNQ mass fraction (1 % ~ 50 %) under simulated sunlight irradiation.



Figure S3 Photocatalytic degradation of 2,4-dichlorophenol (a) and bisphenol A (b), (inset) the apparent rate constants over pure $g-C_3N_4$ (0%) and 10%-TCNQ-C₃N₄ under visible light irradiation (λ >420 nm).



Figure S4 HPLC chromatograms of phenol and after photocatalytic degradation by TCNQ-C₃N₄ (a) and pure C₃N₄ for 4 h monitored at 275 nm ([phenol] = 5 ppm, catalyst = 25 mg/50 mL)



Figure S5 IR spectra of g-C₃N₄, pure TCNQ and TCNQ-C₃N₄ materials.



Figure S6 Mott-Schottky (MS) plots of pure C_3N_4 film electrodes at a frequency of 10 Hz and 100 Hz in an aqueous solution of Na_2SO_4 (0.1 M).



Figure S7 Mott-Schottky (MS) plots of pure TCNQ film electrodes at a frequency of 10 Hz and 100 Hz in an aqueous solution of Na_2SO_4 (0.1 M).



Figure S8 The valence band spectra of X-ray photoelectron spectroscopy for pure C_3N_4 , 10%-TCNQ- C_3N_4 and pure TCNQ.



Figure S9 ESR spectra of 10%-TCNQ-C₃N₄ in dark (a) 10%-TCNQ-C₃N₄ under visible light irradiation ($\lambda > 420$ nm) in water (b) pure C₃N₄ in dark (c) and pure C₃N₄ under visible light irradiation ($\lambda > 420$ nm) in water (d). \bigstar label as superoxide radicals \blacklozenge label as hydroxy radical.



Figure S10 The plots of photogenerated carriers trapping in the system of photodegradation of phenol by 10%-TCNQ-C₃N₄ and pure C₃N₄ under visible light irradiation ($\lambda > 420$ nm)



Figure S11 BET specific surface areas of pure $g-C_3N_4$, pure TCNQ and TCNQ- C_3N_4 with different TCNQ mass fraction.