

Electronic Supplementary Information

Hierarchical architectures of monodisperse porous Cu microspheres: synthesis, growth mechanism, high-efficiency and recyclable catalytic performance

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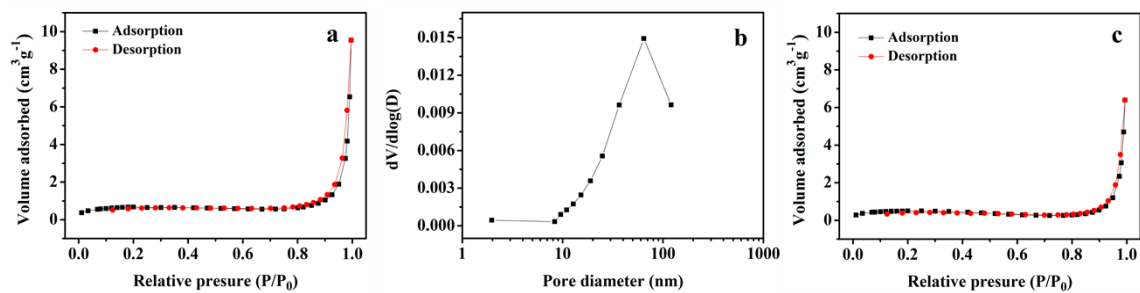


Fig. S1. N₂ adsorption/desorption isotherm curve (a) and the pore sizes distribution (b) of the synthesized porous Cu microspheres; N₂ adsorption/desorption isotherm curve of the solid Cu spheres (c).

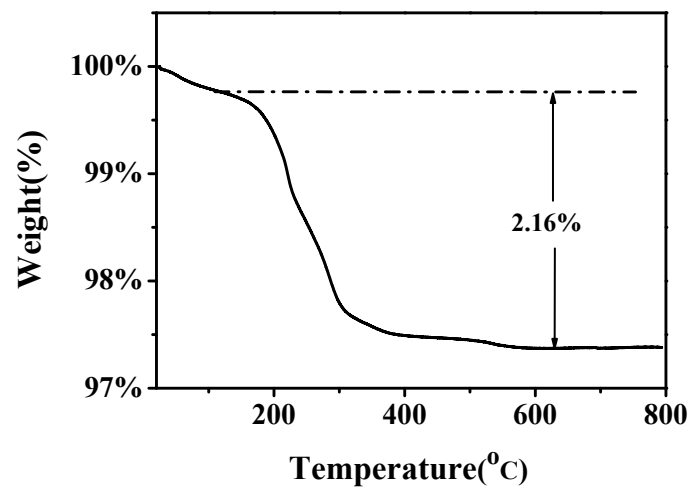


Fig. S2. The TG result of synthesized porous Cu microspheres.

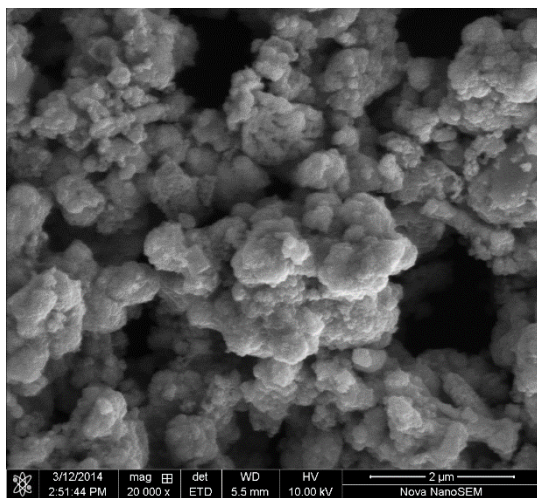


Fig. S3. SEM image of the product with copper chloride as precursor.

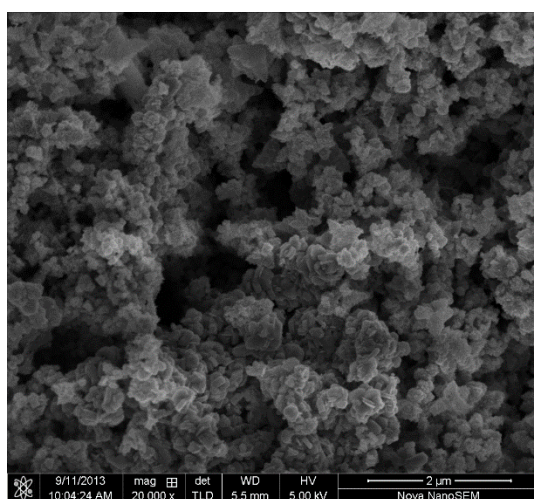


Fig. S4. SEM image of the product without PAA.

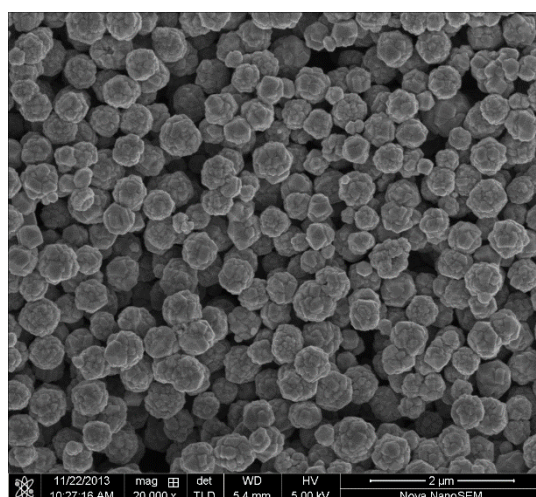


Fig. S5. SEM image of prepared solid copper particles.

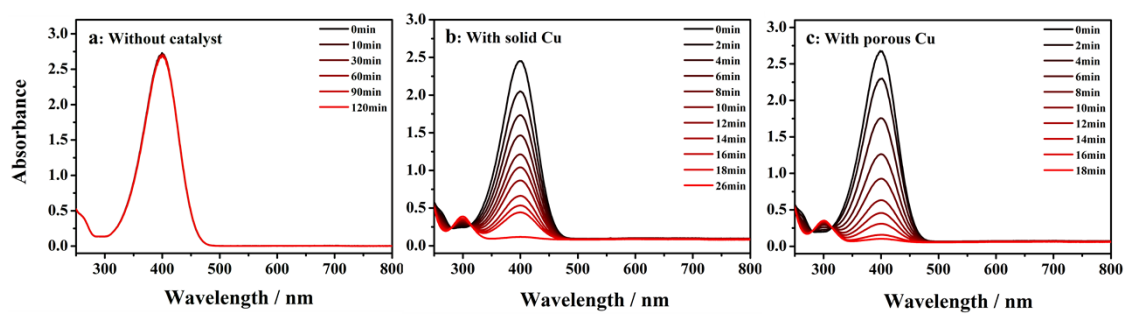


Fig. S6. UV-visible absorption spectra of reduction of 4-NP by NaBH_4 at three different conditions. (a) without catalyst; (b) with solid Cu particles; (c) with porous Cu microspheres.

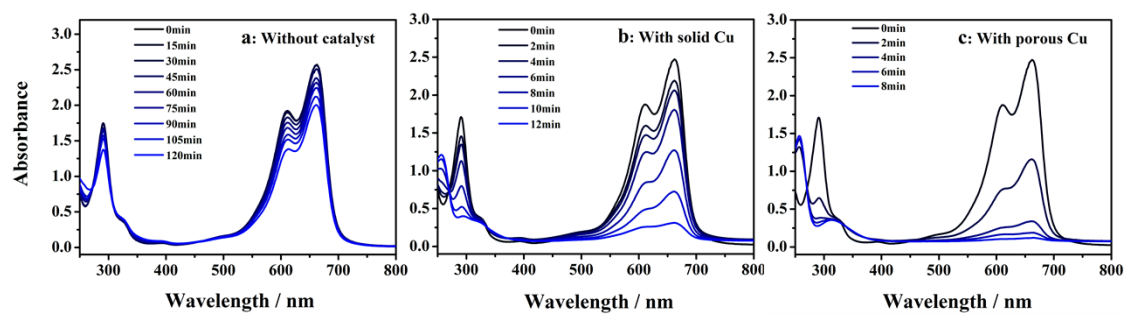


Fig. S7. UV-visible absorption spectra of reduction of MB by NaBH_4 at three different conditions. (a) without catalyst; (b) with solid Cu particles; (c) with porous Cu microspheres.

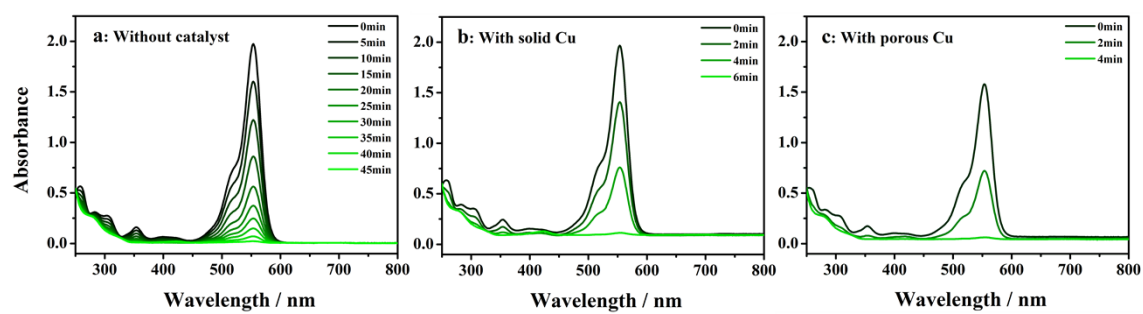


Fig. S8. UV-visible absorption spectra of reduction of RhB by NaBH_4 at three different conditions. (a) without catalyst; (b) with solid Cu particles; (c) with porous Cu microspheres.

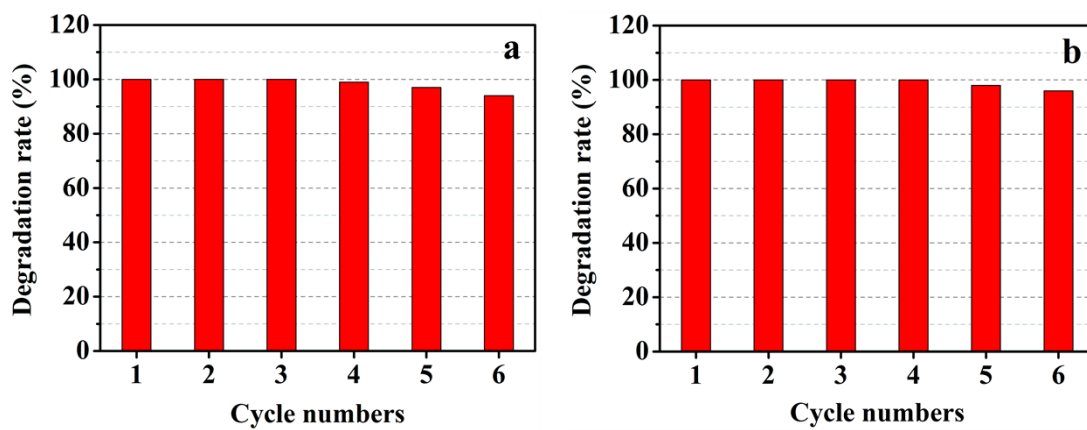


Fig. S9. Stability test of the prepared hierarchical porous Cu microspheres in degradation of MB (a) and RhB (b).