Electronic Supplementary Information for

Synthesis of Self-Stacked $CuFe_2O_4$-Fe_2O_3 Porous Nanosheets as High Performance Li-Ion Battery Anode

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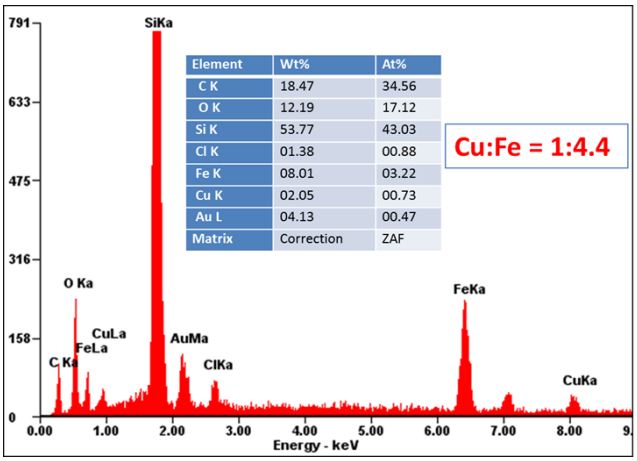

Fig. S1 EDS of as-prepared CuFe-glycolate precursor.

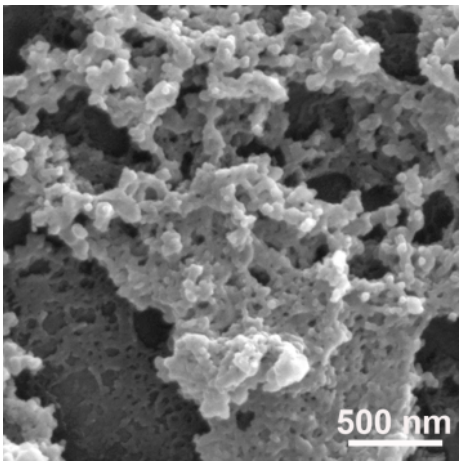

Fig. S2 SEM image of as-prepared Fe-glycolate precursor nanoparticles.

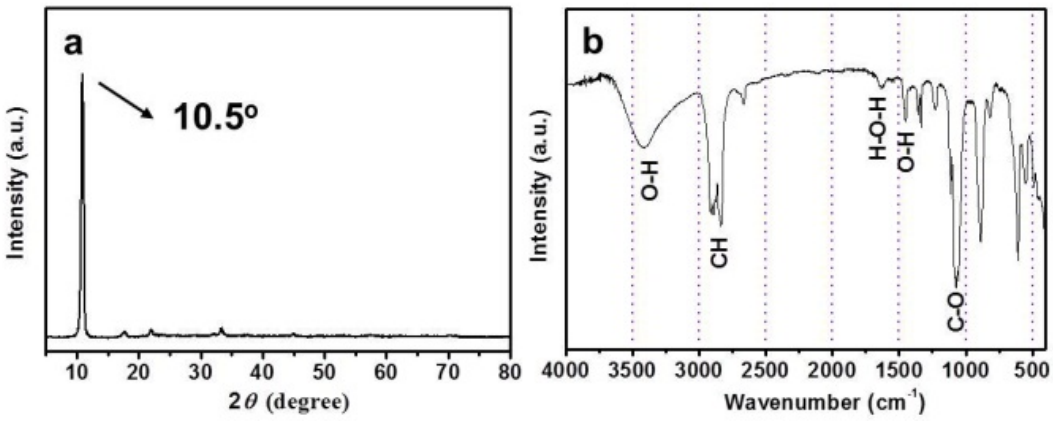

Fig. S3 XRD pattern and FTIR spectra of the as-prepared CuFe-glycolate precursor.

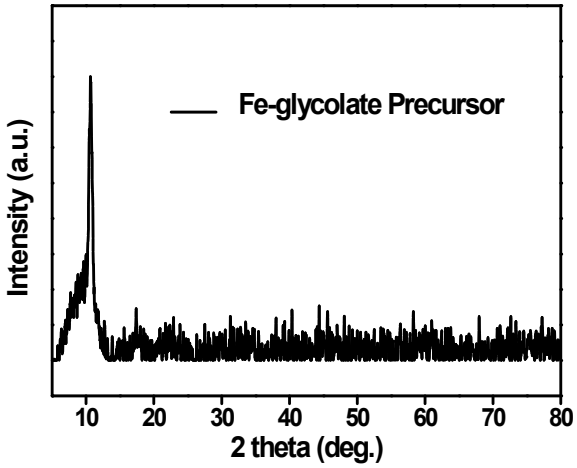

Fig. S4 XRD pattern of as-prepared Fe-glycolate precursor nanoparticles.

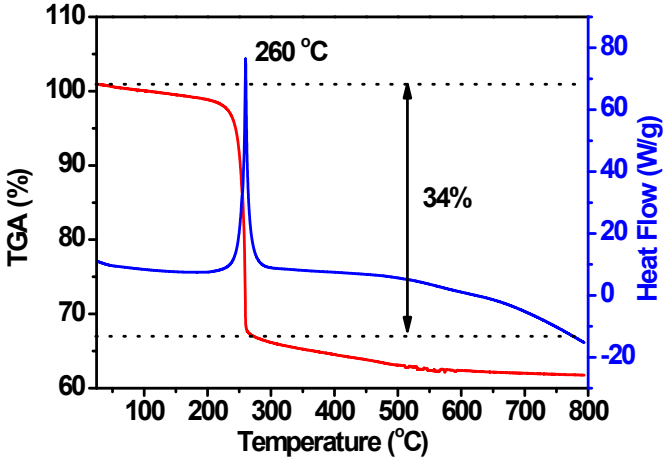

Fig. S5 TGA curve of CuFe-glycolate precursor recorded in air atmosphere.

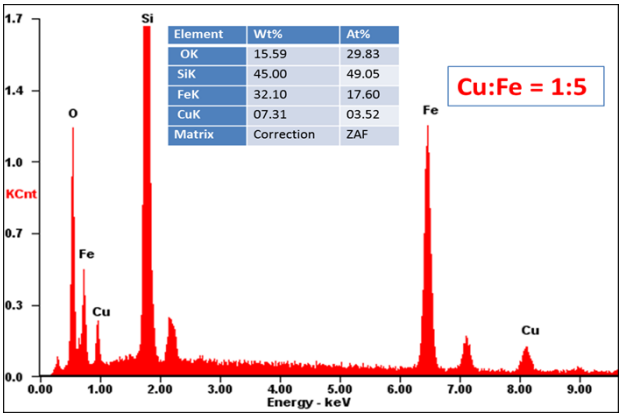

Fig. S6 EDS of the as-synthesized $CuFe_2O_4$-Fe_2O_3 nanosheets after annealing at 350 ºC for 3 h in air.

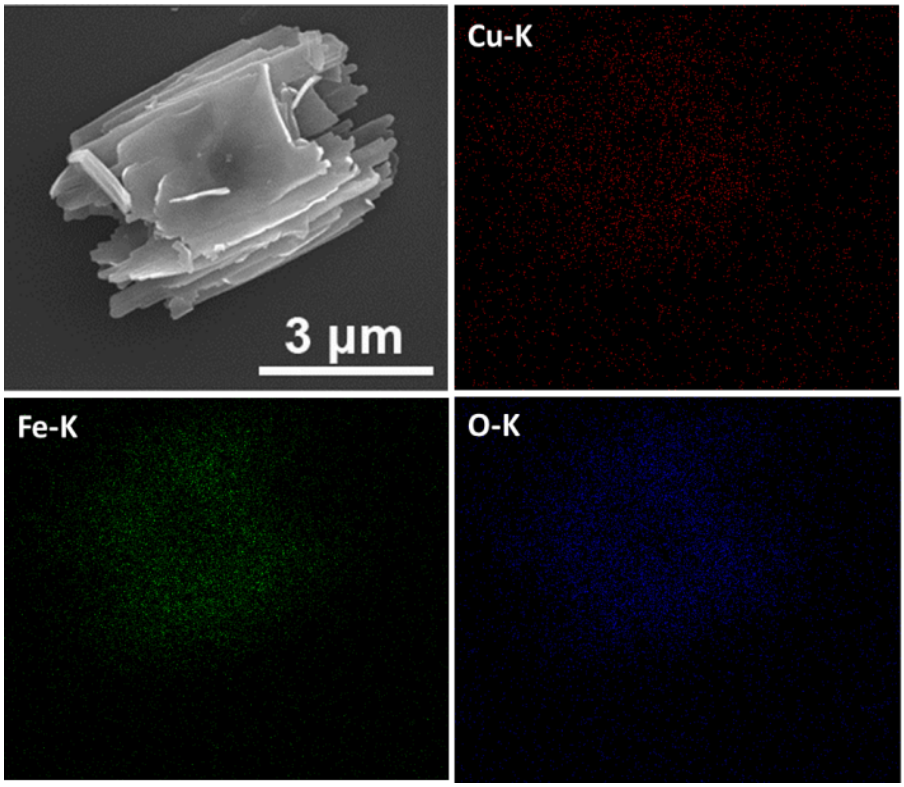

Fig. S7 Element mapping images of Cu, Fe and O of a single self-stacked $CuFe_2O_4$-Fe_2O_3 nanosheets.

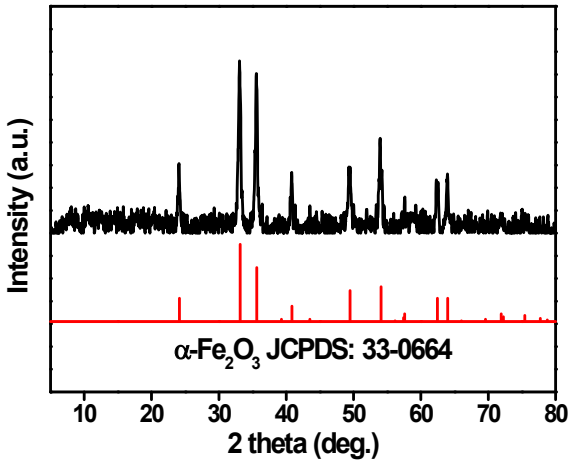

Fig. S8 XRD pattern of the thermal annealing Fe-glycolate precursor.

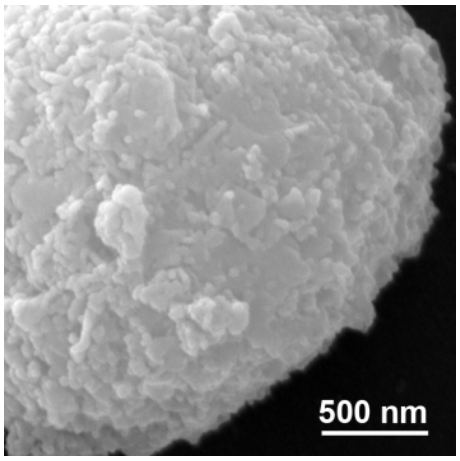

Fig. S9 SEM image of the Fe_2O_3 nanoparticles after heat treatment.

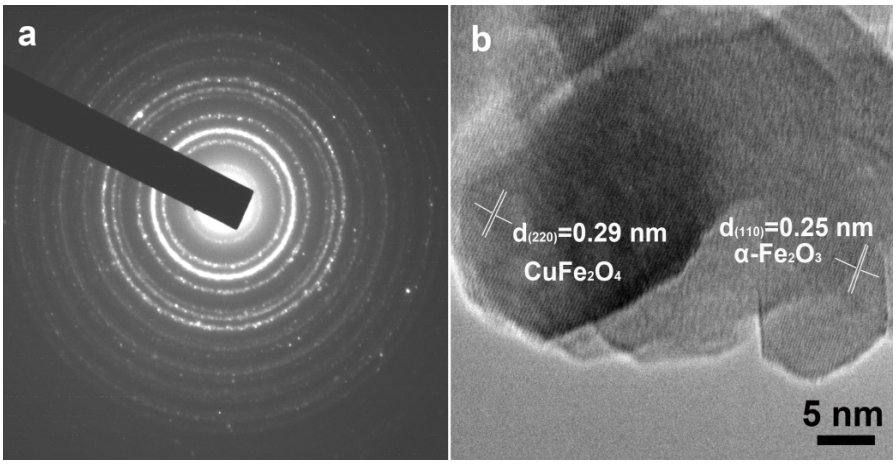

Fig. S10 SAED pattern (a) and HRTEM image (b) of the as-synthesized $CuFe_2O_4$-Fe_2O_3 nanosheets after thermal annealing at 350°C for 3 h in air.

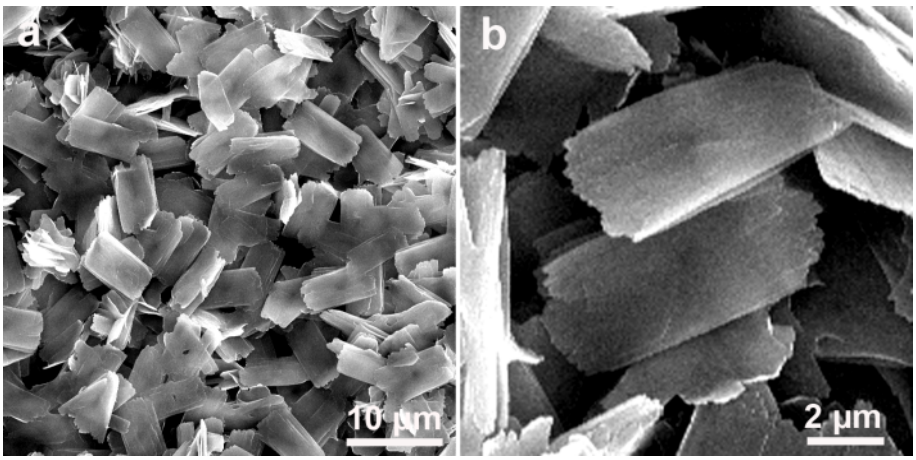

Fig. S11 SEM images of as-prepared CuFe-glycolate precursor nanosheets with a short reaction time of 30 min.

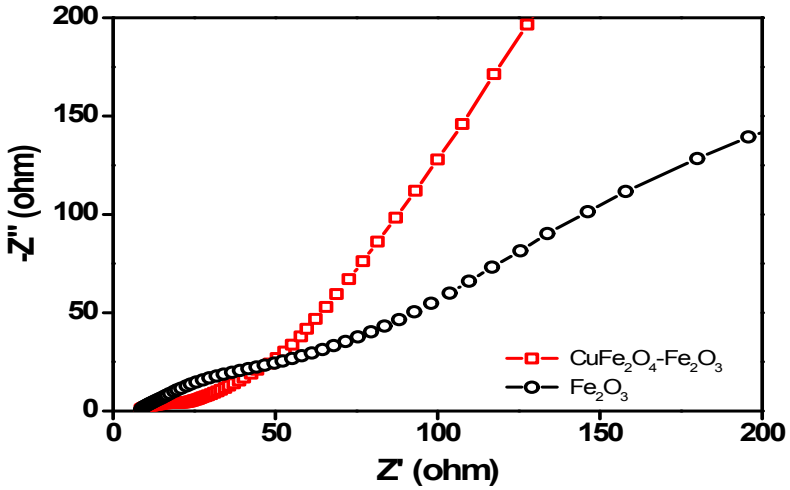

Fig. S12 Nyquist plots of self-stacked $CuFe_2O_4$-Fe_2O_3 nanosheets and Fe_2O_3 nanosheets electrodes after 100 discharge/charge cycles.